

Chemistry Practise Paper

CLASS - IX
SUMMATIVE ASSESSMENT - II, 2013-14
Section - A

TN-M-17

1.

- (1) Which of the following is not soluble in water?
(a) BaCl_2 (b) NaCl (c) BaSO_4 (d) AgNO_3 ①
- (2) While performing experiments to verify the law of Conservation of Mass in a chemical reaction we generally use 5% solution of two salts. 5% salt solution means 5g salt is equal to
(a) 100g of water (b) 95g of water
(c) 105g of water (d) 100 ml. of water ①

Section B

~~(3)~~

- (3) Define atomic mass unit. ①
- (4) Give the chemical formula for the following compounds and compute the ratio by mass of the combining elements in each one of them.
(a) Ammonia (b) Carbon Monoxide (c) Hydrogen Chloride ③
- (5) An atom of an element has one electron in the outermost M shell. state its :
(a) Electronic Configuration (b) Number of protons
(c) Atomic Number (d) valency of the element

(e) Formula of its chloride and sulphate. (3)

(6)

(a) How many atoms are there in 0.5 mole of Cl_2 molecules?

(b) Sample A contains one gram of mole of oxygen molecules and Sample B contains one mole of oxygen molecules? what is the ratio of the number of molecules in both the samples? (3)

(7) Give reasons for the following:

(a) Isotopes of an element are chemically similar.

(b) An atom is electrically neutral.

(c) Noble gases show least reactivity.

(d) Nucleus of an atom is heavy and positively charged.

(e) Ions are more stable than atoms. (5)